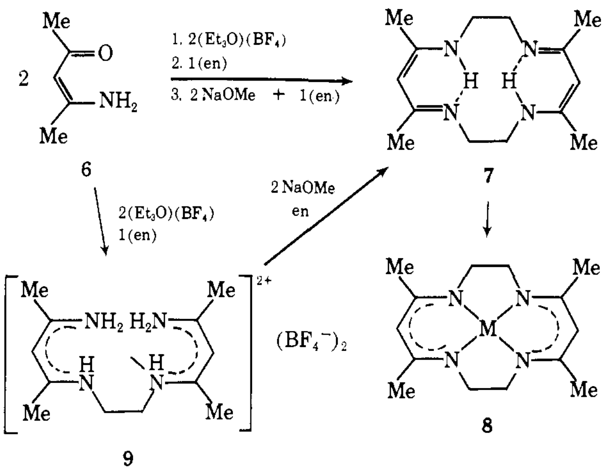
−11.61 ppm (2, NH); mol wt 244 (toluene, osmometry), calcd 248; mass spectrum (70 eV) m/e 248 (P). Nickel(II) and copper(II) complexes (**8**) were prepared from **7** and the metal acetate in ethanol, while the remaining complexes were obtained by nonaqueous chelation reactions in *tert*-butyl alcohol[10] (Fe(II), Co(II)) or THF[11,12] (Zn(II)). As a class the complexes M(MeHMe(en)$_2$) are easily crystallized and moderately soluble in weakly polar solvents. As solids, all are stable to dry oxygen except the Fe(II) species; this complex and Co(MeHMe(en)$_2$) react with oxygen in solution. The spectral and magnetic properties are consistent with a planar stereochemistry.[13]



The reaction sequence in the conversion **6** → **7** is under further investigation. Interruption of the reaction after addition of the first equivalent of ethylenediamine has allowed isolation of the intermediate salt **9**, the exact tautomeric structure of which is uncertain. This material was purified by recrystallization from dry methanol and isolated as a very hygroscopic white solid (34%, mp 175–180°). Neutralization of **9** followed by reaction with 1 equiv of ethylenediamine in dry methanol gives the macrocycle **7** in 50% yield. The use of **9** as an intermediate in the preparation of unsymmetrical macrocycles is under study. The synthetic scheme described here represents a particularly clear example of the utility of *O*-alkyl-β-ketoamine cations as intermediates in the synthesis of open[12,14] or cyclic ligand systems by nucleophilic reactions at the >COR carbon, and should be suggestive of further applications.

Details of the electronic and redox properties and the oxidative reactivities of M(MeHMe(en)$_2$) and related complexes, as well as additional experiments directed toward a useful template synthesis of **7** and **8**, will be reported subsequently.

Acknowledgment. This research was supported by grants from the National Institutes of Health (No. GM-15471) and the National Science Foundation (No. GP-7576X).

T. J. Truex, R. H. Holm*
Department of Chemistry
Massachusetts Institute of Technology
Cambridge, Massachusetts 02139

Received October 12, 1970

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[12] D. H. Gerlach and R. H. Holm, *ibid.*, **91**, 3457 (1969).

[13] Magnetic moments of crystalline solids (**8**) at ~25°: M = Ni(II), Zn(II) diamagnetic; Cu(II), 1.78 BM; Co(II), 2.08 BM; Fe(II), 3.95 BM.

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Additions and Corrections

Kinetics and Mechanism of the Osmium Tetroxide Catalyzed Oxidation of Acetone and Ethyl Methyl Ketone by Alkaline Hexacyanoferrate(III) Ion [*J. Amer. Chem. Soc.*, **91**, 2643 (1969)]. By V. N. Singh, H. S. Singh, and B. B. L. Saxena, Department of Chemistry, University of Allahabad, Allahabad, India.

The right-hand side of eq 1–4, 6, and 7 should be multiplied by 2 and that of eq 5 by $^1/_2$. The resulting rate expressions thus would be valid for the first stage of oxidation. The k_1 values calculated should be read as $2k_1$. The species of osmium(VIII) and osmium(VI) should be taken as $OsO_4(OH)_2^{2-}$ and $OsO_2(OH)_4^{2-}$.

On the Mechanism of Interaction between Tertiary Amines and Trichlorosilane [*J. Amer. Chem. Soc.*, **92**, 699 (1970)]. By Stanley C. Bernstein, Department of Chemistry, Wright State University, Dayton, Ohio 45431

The rate equation on page 699 should read

$$4AIk_{FD}t = 2(I-1)(x-1) - (I+1)\ln(2x-1)$$

The calculations in the communication are based on this correct equation.

Conformational Analysis. XXII. Conformational Equilibria in 2-Substituted 1,3-Dioxanes [*J. Amer. Chem. Soc.*, **92**, 3050 (1970)]. By Franz W. Nader and Ernest L. Eliel, Department of Chemistry, University of Notre Dame, Notre Dame, Indiana 46556.

In Table VI, column 1, compound II should read III and III should read IV. The first two footnotes to the table should read: ^{a} Axial isomer/equatorial isomer. ^{b} Equatorial isomer/axial isomer.

A Model for the Biogenesis of the Spirobenzylisoquinoline Alkaloids [*J. Amer. Chem. Soc.*, **92**, 4943 (1970)]. By